

Chapter 12-5: Gas Molecules

- Avogadro's hypothesis
 - Determine the volume occupied by 0.202 moles of a gas at STP.
 - How many oxygen molecules are in 3.36g of oxygen gas at STP?
 - What is the volume of 14.0g of nitrogen gas at STP?
- Dalton's Law of Partial Pressures
 - Practice Problem. Air contains several gases mixed together. IF the partial pressure at 101.3 kPa is made up of nitrogen at 79.1 kPa, carbon dioxide at 0.040 kPa, and some assorted gases at 0.94kPa, how much partial pressure will oxygen have?
- Graham's Law of Effusion

$$\frac{Rate_A}{Rate_B} = \frac{\sqrt{mm_B}}{\sqrt{mm_A}}$$

Homework: Section 12-5 (p353) #31-36, (p351) #37-38 (p353) #39-44