

Chapter 11-1: Heat and the Flow of Energy

- Thermochemistry
- Energy
- Chemical Potential Energy
- Heat

- Δ
- Exothermic and Endothermic Processes
 - System versus Surroundings
 - Law of Conservation of Energy
 - Endothermic:
 - Exothermic:
- Heat Capacity
 - **Calories**, kilocalories, calories, Joules

 - How many **Calories** are contained in 1 Joule?

 - Specific Heat

one must put into a substance

$$q = m \cdot \Delta t \cdot C$$

STUDENTS: Please copy Table 11.2 on page 296 into the notes here

- How many Joules are required to heat 21.7g of water from 18.2°C to 39.7°C?

- What is the final temperature after 777J of energy is added to 13.9g of water at 21.9°C?

- What is the final temperature after 777J of energy is added to 13.9g of silver at 21.9°C?

Homework: Section 11-1 (p299) #1-10