

Chapter 9-3: Limiting Reagent and Percent Yield

- Limiting Reagent
 - Excess Reagent
- Determining the Limiting Reagent

- Percent Yield

$$\%Yield = \frac{\text{actual}}{\text{theoretical}} \times 100\%$$

- EX: $2\text{H}_2\text{S}_{(g)} + 3\text{O}_{2(g)} \rightarrow 2\text{SO}_{2(g)} + 2\text{H}_2\text{O}_{(g)}$
If given 78g H_2S and 100g O_2 , which is the limiting reagent?
How much of the other will be left?
What is the highest yield you can obtain of SO_2 ?
If you actually obtain 100g of SO_2 , what is your percent yield?

Homework: Section 9-3 (p259) #29-32